

Please check the examination details below before entering your candidate information

Candidate surname		Other names	
Centre Number		Candidate Number	

Pearson Edexcel Level 1/Level 2 GCSE (9–1)

Monday 19 May 2025

Morning (Time: 1 hour 45 minutes)

Paper reference **1CH0/1H**

Chemistry

PAPER 1

Higher Tier

You must have:
Calculator, ruler, Periodic Table (enclosed)

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*
- Calculators may be used.
- Any diagrams may NOT be accurately drawn, unless otherwise indicated.
- You must **show all your working out** with **your answer clearly identified** at the **end of your solution**.

Information

- The total mark for this paper is 100.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*
- In questions marked with an **asterisk** (*), marks will be awarded for your ability to structure your answer logically, showing how the points that you make are related or follow on from each other where appropriate.

Advice

- Read each question carefully before you start to answer it.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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Pearson

Answer ALL questions. Write your answers in the spaces provided.

Some questions must be answered with a cross in a box ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

- 1 (a) Figure 1 shows a chemical cell.

The chemical cell contains two metal electrodes placed into an electrolyte and connected in a circuit with a voltmeter.

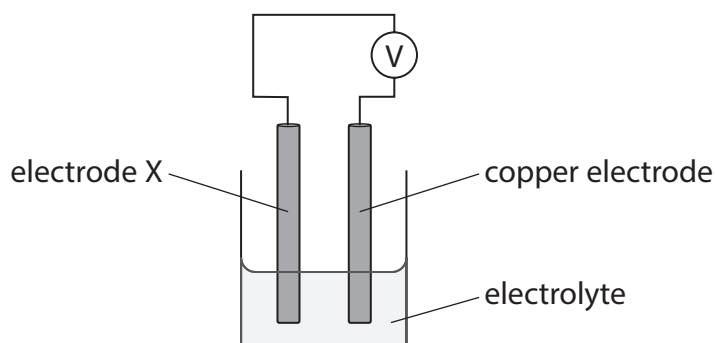


Figure 1

- (i) State what type of compound must be present in the electrolyte.

(1)

- (ii) When the two metal electrodes are placed in the electrolyte, a voltage is produced.

The bigger the difference in reactivity of the metals, the larger the voltage produced.

Which metal, when used for electrode X in Figure 1, would give the largest voltage?

(1)

- ☐ A copper
- ☐ B gold
- ☐ C magnesium
- ☐ D zinc



(iii) State what happens to the voltage when one of the reactants is used up.

(1)

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(b) Hydrogen-oxygen fuel cells can also be used to produce a voltage.

(i) Draw the dot and cross diagram for a molecule of oxygen, O_2 .

An atom of oxygen has the electronic configuration of 2.6.

Show outer electrons only.

(2)

(ii) Hydrogen-oxygen fuel cells are used in spacecraft.

Suggest **one advantage** of using fuel cells rather than chemical cells in spacecraft.

(1)

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(Total for Question 1 = 6 marks)

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2 (a) Atoms of neon are made up of subatomic particles.

- (i) Figure 2 shows some information about the subatomic particles in neon atoms.

Complete Figure 2 with the missing information.

(3)

subatomic particle	relative charge	relative mass
proton		1
neutron	0	
electron		negligible

Figure 2

- (ii) All atoms of neon have an atomic number of 10.

Describe what information about the number of subatomic particles can be obtained from the atomic number of neon.

(2)

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- (iii) Explain why atoms of neon are electrically neutral.

(2)

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(b) A sample of neon contains three isotopes

90.48% of neon-20

0.27% of neon-21

9.25% of neon-22

Using this information, calculate the relative atomic mass of this sample of neon.

(2)

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relative atomic mass =

(Total for Question 2 = 9 marks)



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3 Silver nitrate and potassium chloride are ionic compounds that are soluble in water.

(a) When potassium chloride solution is added to a solution of silver nitrate, a white precipitate forms.

(i) Complete the equation for this reaction, including the missing state symbol.

(2)



(ii) Give the name of the white precipitate that forms.

(1)

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(iii) State and explain how the white precipitate can be separated from the reaction mixture.

(3)

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(b) Silver nitrate solution conducts electricity.
Solid silver also conducts electricity.

Describe how silver nitrate solution conducts electricity and how solid silver conducts electricity.

(3)

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(Total for Question 3 = 9 marks)



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4 Phosphoric acid is important in the production of fertilisers.

- (a) (i) A few drops of litmus solution are added to a sample of dilute phosphoric acid.

What is the colour of the mixture formed?

(1)

- ☐ **A** colourless
- ☐ **B** green
- ☐ **C** purple
- ☐ **D** red

- (ii) State why litmus cannot be used to measure pH.

(1)

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- (b) Phosphoric acid can react to form phosphate salts which are used in fertilisers.

Give **two** reasons why farmers use fertilisers on their fields.

(2)

1

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2

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(c) An experiment was carried out to produce phosphoric acid.

(i) The actual yield for the experiment was 18.62 g.

The theoretical yield for the experiment is 24.50 g.

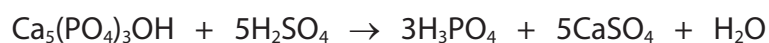
$$\text{percentage yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

Calculate the percentage yield of this experiment.

(2)

percentage yield =

(ii) The equation for a reaction to make phosphoric acid, H_3PO_4 , is



Calculate the atom economy for the production of phosphoric acid in this reaction.

$$\text{atom economy (\%)} = \frac{\text{total formula mass of desired product}}{\text{total formula mass of all reactants or products}} \times 100$$

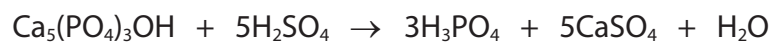
(relative formula masses: $\text{Ca}_5(\text{PO}_4)_3\text{OH} = 502$, $\text{H}_2\text{SO}_4 = 98$, $\text{H}_3\text{PO}_4 = 98$, $\text{CaSO}_4 = 136$, $\text{H}_2\text{O} = 18$)

(3)

atom economy = %



(iii) The equation for a reaction to produce phosphoric acid is



Explain how the atom economy of the reaction can be increased.

(2)

(Total for Question 4 = 11 marks)

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5 Nitrogen, N_2 , oxygen, O_2 and fluorine F_2 are all gases.

(a) Nitrogen, oxygen and fluorine are in the same period of the periodic table.

Explain, in terms of electronic configuration, in which period of the periodic table nitrogen, oxygen and fluorine are located.

(2)

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(b) Fluorine has a boiling point of -188°C .

(i) Explain why fluorine has a low boiling point.

(3)

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(ii) A mixture of nitrogen and oxygen becomes a liquid when cooled to -200°C .

Which row shows the changes in energy and movement of the molecules as they are cooled from room temperature to -200°C ?

(1)

	movement of molecules	energy of molecules
☐ A	less	more
☐ B	more	less
☐ C	less	less
☐ D	more	more



(iii) The mixture of liquid nitrogen and liquid oxygen can be separated.

Which process could be used to separate a mixture of liquid nitrogen and liquid oxygen?

(1)

- ☐ A paper chromatography
- ☐ B fractional distillation
- ☐ C filtration
- ☐ D crystallisation

(c) At high temperatures, nitrogen, N_2 , can react with oxygen, O_2 , to form nitrogen dioxide only.

Write the balanced equation for this reaction.

(2)

(Total for Question 5 = 9 marks)

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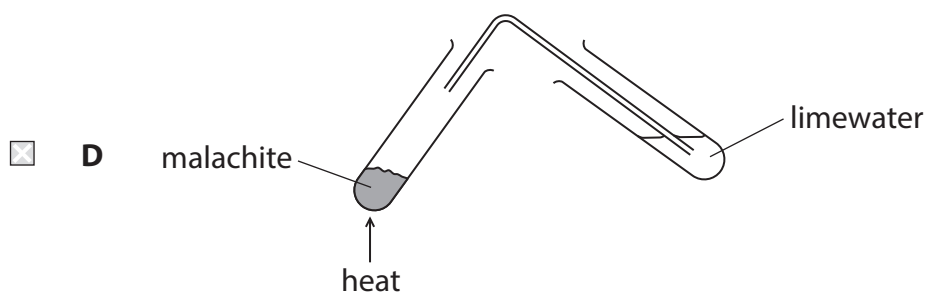
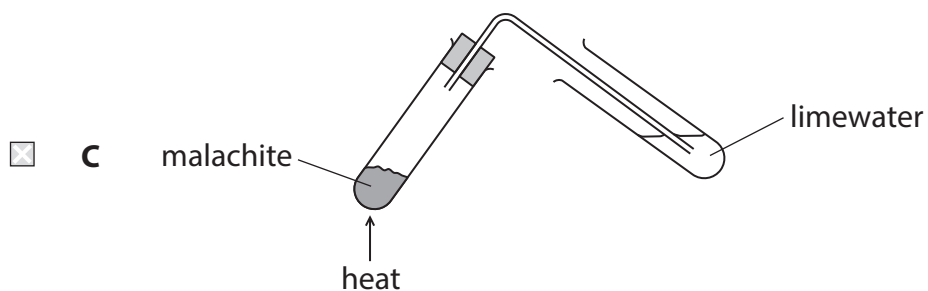
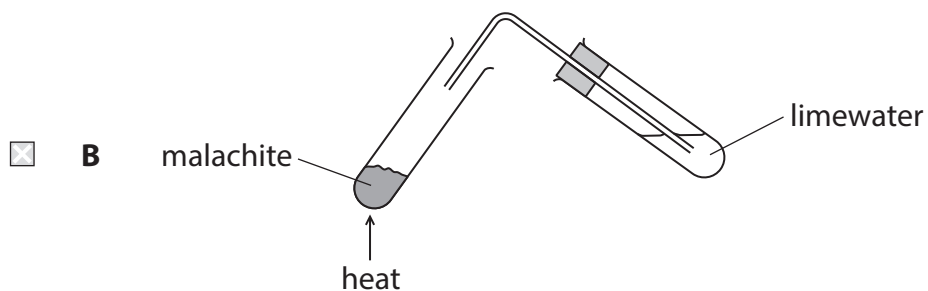
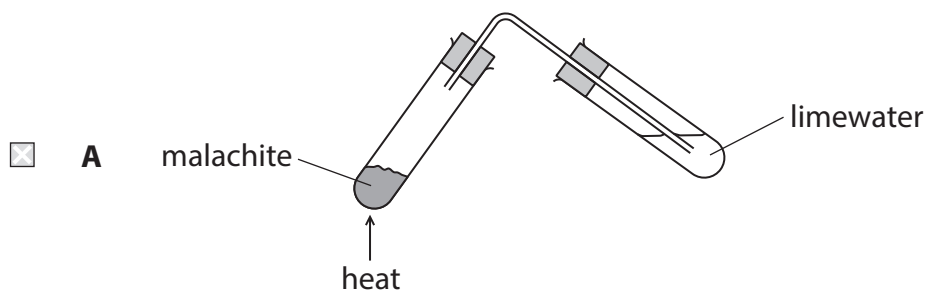


6 Malachite, $\text{Cu}_2\text{CO}_3(\text{OH})_2$, is a mineral containing copper.

(a) When heated, malachite decomposes to form copper oxide, carbon dioxide and water.

(i) Which apparatus should be used to heat the malachite and test for carbon dioxide?

(1)

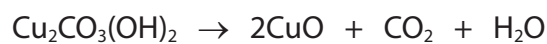


- (ii) When carbon dioxide is bubbled through limewater, a precipitate forms.
This precipitate solid makes the limewater cloudy.

Name the precipitate that makes the limewater cloudy.

(1)

- (iii) The equation for the decomposition of malachite when it is heated is



Calculate the maximum mass of copper oxide that can be formed when
22.1 g of malachite is decomposed completely.

(relative atomic masses: Cu = 63.5, O = 16.0, C = 12.0, H = 1.0)

(4)

maximum mass of CuO = g

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- (b) Copper can be extracted from copper oxide by heating the copper oxide with carbon.

A student investigated two methods to extract the maximum mass of copper from copper oxide.

Figure 3 shows the two methods.

method 1	method 2
step 1 place one spatula measure of copper oxide into a beaker	step 1 place one spatula measure of copper oxide into a glass test tube
step 2 mix the copper oxide thoroughly with carbon powder	step 2 carefully add carbon powder on top of the copper oxide
step 3 place the mixture into a metal crucible and heat the mixture with a roaring Bunsen flame	step 3 heat the two layers with a roaring Bunsen flame

Figure 3

Comment on the strengths and weaknesses of these two methods.

(3)

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(Total for Question 6 = 9 marks)

7 Copper, iron, nickel and zinc are all metallic elements.

(a) Which row of the table applies to copper?

(1)

	type of metal	property
☐ A	it is an alkali metal	it has a low melting point
☐ B	it is an alkali metal	it forms coloured compounds
☐ C	it is a transition metal	it has a high melting point
☐ D	it is a transition metal	it forms coloured compounds

(b) Many ships are made of steel.

Steel is an alloy that contains iron.

Figure 4 shows blocks of zinc attached to a steel ship.

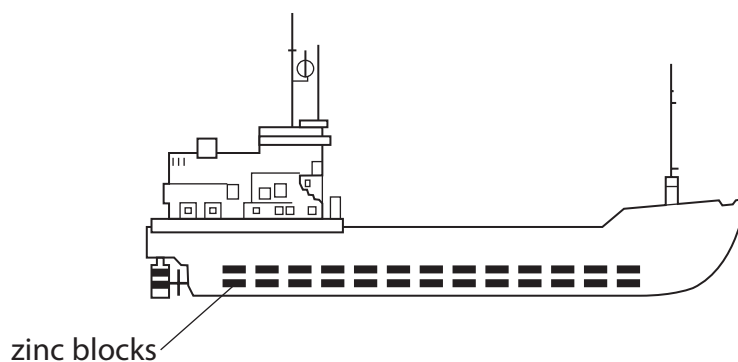


Figure 4

Explain how iron is protected from rusting by the zinc blocks.

(3)

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(c) Alloys of copper and nickel are used for some coins.

Figure 5 shows the relative strengths of copper and nickel and some of their alloys.

percentage of nickel	relative strength
0	1.00
20	1.52
35	1.77
65	1.90
80	1.78
100	1.56

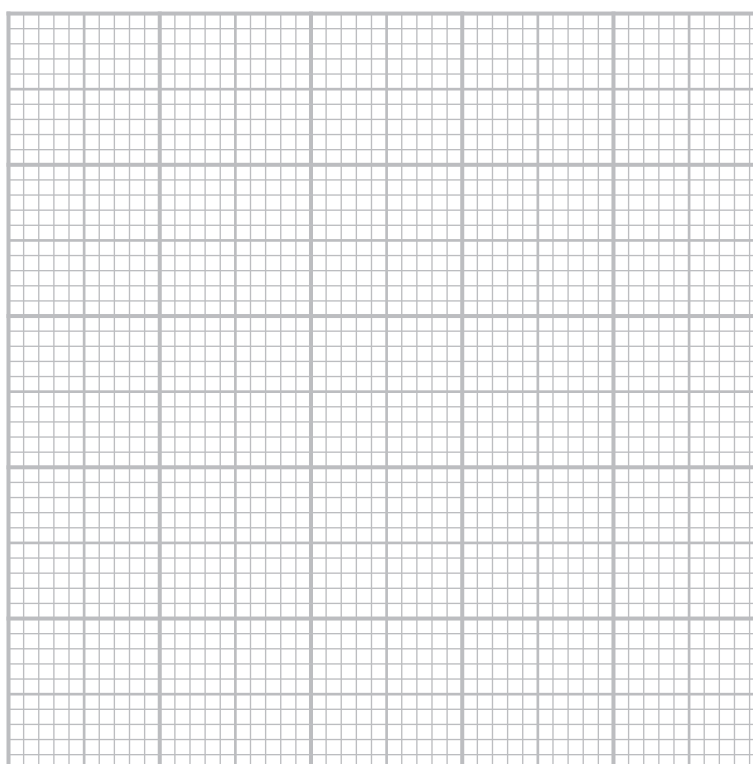
Figure 5

(i) On Figure 6

- add suitable scales to the vertical and horizontal axes
- plot a graph of relative strength against percentage of nickel
- draw a curve of best fit.

(4)

relative strength



percentage of nickel

Figure 6



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(iii) Explain, using Figure 6, how the percentage of nickel affects the relative strength of the alloys of copper and nickel.

(Total for Question 7 = 12 marks)



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8 Magnesium sulfate is an ionic compound.

(a) A student adds magnesium sulfate to water to form a solution.

Which of these will also form a solution when added to water?

(1)

- ☐ **A** barium sulfate
- ☐ **B** calcium sulfate
- ☐ **C** lead sulfate
- ☐ **D** sodium sulfate

(b) (i) Calculate the percentage by mass of oxygen in magnesium sulfate, MgSO_4 .

(relative atomic mass: $\text{O} = 16.0$

relative formula mass: $\text{MgSO}_4 = 120$)

Give your answer to 3 significant figures.

(3)

percentage by mass =

(ii) Calculate the number of oxygen atoms in 7.20 g of MgSO_4 .

(relative formula mass: $\text{MgSO}_4 = 120$

Avogadro constant = 6.02×10^{23})

(3)

number of oxygen atoms =

*(c) Magnesium sulfate solution can be electrolysed using the apparatus shown in Figure 7.

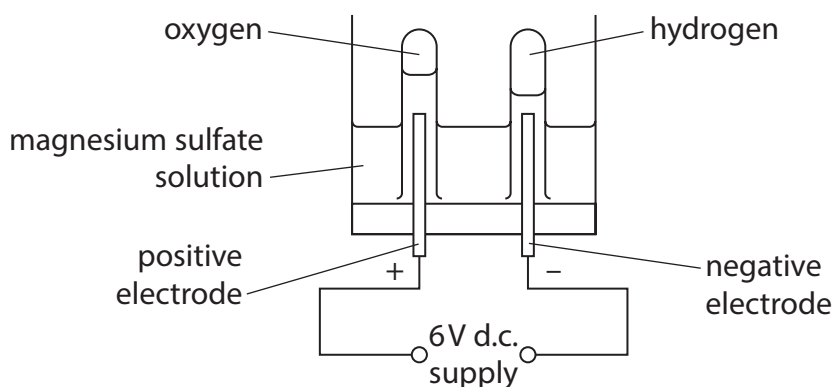


Figure 7

Explain how the products are formed at the electrodes in terms of

- the ions present in the solution
- the oxidation and reduction reactions that take place at the electrodes.

You should include ionic equations as part of your answer.

(6)

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Handwriting practice area with 20 horizontal dotted lines.

(Total for Question 8 = 13 marks)



9 This question is about ammonia.

(a) Ammonia is a gas.

Suggest why the litmus paper used to test for ammonia gas has to be damp.

(2)

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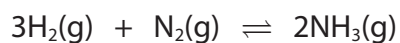
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(b) Ammonia can be made on an industrial scale using the Haber process.

The equation for the reaction is



The reaction can reach a dynamic equilibrium.

(i) Explain what happens when the reaction reaches dynamic equilibrium

(2)

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- (ii) Figure 8 shows the changes in percentage yield of ammonia at equilibrium as temperature increases, at a pressure of 300 atmospheres and at a pressure of 1000 atmospheres.

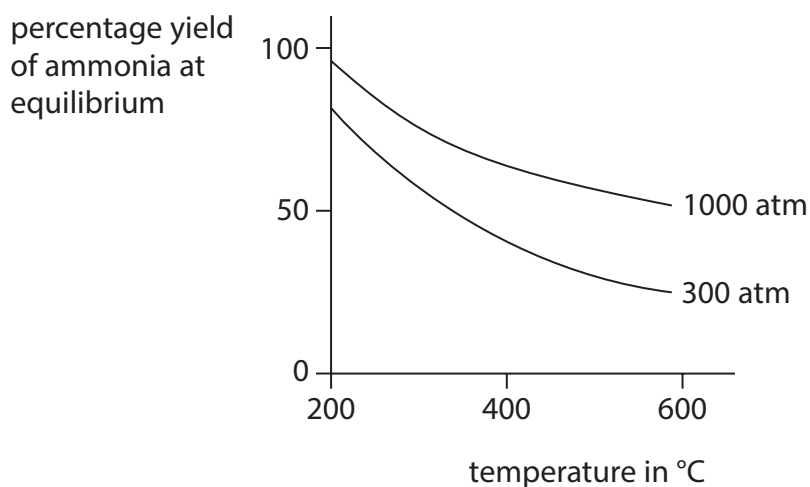


Figure 8

Which row of the table will give the highest percentage yield of ammonia at equilibrium?

(1)

	temperature in °C	pressure in atm
☐ A	200	1000
☐ B	300	1000
☐ C	600	300
☐ D	600	1000

- (iii) In industry, a pressure of 200 atmospheres is used for the Haber process.

Using Figure 8, explain what happens to the position of dynamic equilibrium if the process is carried out at 100 atmospheres.

(2)

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(c) Ammonia can also be made by heating ammonium chloride.



A sample of solid ammonium chloride is heated in a boiling tube.

After a few minutes ammonium chloride, a white solid, is also seen at the top of the boiling tube, as shown in Figure 9.

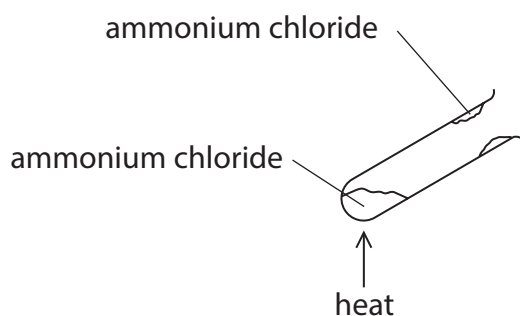


Figure 9

Explain how ammonium chloride is formed at the top of the boiling tube.

(4)

(Total for Question 9 = 11 marks)

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10 A student is asked to carry out a titration to find the concentration of a sodium hydroxide solution.

(a) A teacher dilutes some sulfuric acid with distilled water to make a solution with a concentration of $0.250 \text{ mol dm}^{-3}$ for the student to use.

(i) What effect does diluting the sulfuric acid with distilled water have on the pH of the solution and the concentration of the H^+ ions?

(1)

	concentration of the H^+ ions	pH of the solution
☐ A	decreased	decreased
☐ B	increased	increased
☐ C	increased	decreased
☐ D	decreased	increased

(ii) What is the concentration, in g dm^{-3} , of a $0.250 \text{ mol dm}^{-3}$ solution of H_2SO_4 ?

(relative formula mass: $\text{H}_2\text{SO}_4 = 98.0$)

(1)

- ☐ A 0.00245
- ☐ B 0.394
- ☐ C 24.5
- ☐ D 394



(b) The student uses the apparatus shown in Figure 10 to carry out the titration.

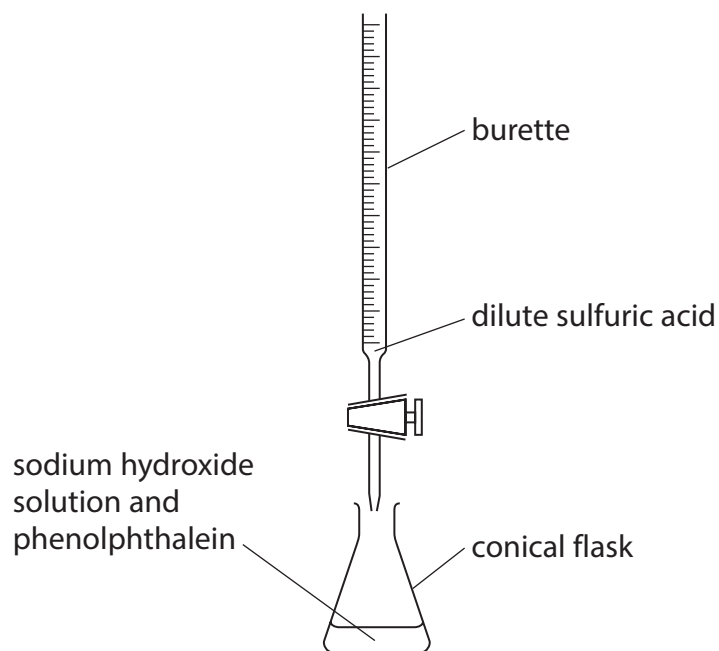


Figure 10

Before using the apparatus, the student rinses the burette with distilled water.

This is an error.

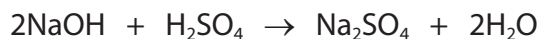
(i) Explain what effect this error has on the result of the titration.

(2)

(ii) State what the student should do to correct the error.

(1)

*(c) The equation for the reaction in the titration is



Explain how the equation for the reaction and the apparatus in Figure 10 can be used to find the concentration of 25.0 cm^3 of sodium hydroxide solution using the 0.250 mol dm^{-3} sulfuric acid.

(6)



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(Total for Question 10 = 11 marks)

TOTAL FOR PAPER = 100 MARKS



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Monday 19 May 2025

Paper
reference

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Chemistry

PAPER 1

Higher Tier

Periodic Table Insert

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The periodic table of the elements

1	2	Key										3	4	5	6	7	0
		relative atomic mass atomic symbol name atomic (proton) number															
7 Li lithium 3	9 Be beryllium 4											11 B boron 5	12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10
23 Na sodium 11	24 Mg magnesium 12											27 Al aluminium 13	28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18
39 K potassium 19	40 Ca calcium 20	45 Sc scandium 21	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26	59 Ni nickel 28	63.5 Cu copper 29	65 Zn zinc 30	70 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	80 Br bromine 35	84 Kr krypton 36	
85 Rb rubidium 37	88 Sr strontium 38	89 Y yttrium 39	91 Zr zirconium 40	93 Nb niobium 41	96 Mo molybdenum 42	[98] Tc technetium 43	101 Ru ruthenium 44	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	128 Te tellurium 52	127 I iodine 53	131 Xe xenon 54	
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86	

1
H
hydrogen
1

Key

relative atomic mass
atomic symbol
name
atomic (proton) number

* The elements with atomic numbers from 58 to 71 are omitted from this part of the periodic table.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.